BIG ChEM

BLITZ Unit 7

PRINT NAME _____ PERIOD_____

*** You MUST USE INK. You may use your notes, but no help from others.

*** BE SURE TO BALANCE YOUR EQUATIONS!**
*** SHOW ALL STEPS**
*** WATCH FOR DIATOMIC ELEMENTS.**

1. Find the number of grams of ammonia (NH_3) that are formed from 2.00 g of hydrogen (H_2).

2. Using $Q = mc\Delta t$, find how much heat will raise the temperature of 4.66 g of CCl_4 from 20.9°C to 76°C. From the table, c for CCl_4 is 0.856 J/g.C°.

3. Find ΔH for making 193 g of NH_4Br from NH_3 and HBr. From the table, the ΔH's in kJ/mol: NH_4Br = -270, NH_3 = -46.2, HBr = -36.2

4. DEFINE and GIVE an example of these:

   endothermic reaction
   exothermic reaction
   activation energy

5. DRAW and LABEL the curves for Endothermic and Exothermic reactions.